



## FIB ANSWERS:

1. Periods
2. Decreases
3. Increases
4. Metallic
5. Atomic numbers
6. Period, non-metallic
7. More
8. Number of outer electrons
9. Decreases, magnitude of nuclear charge, number of shells
10. Metal
11. Larger
12. More
13. C
14. Lesser
15. Metallic
16. Smallest
17. Ionic
18. Bottom
19. Atomic number

20. Non metal
21. Sixth
22. Halogens
23. Decreases
24. Aluminium oxide
25. Bromine
26. Less
27. Lower
28. Sodium
29. Chlorine

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30. Cathode, anode
  31. Three single
  32. Three, one
  33. Solid
  34. Low
  35. Loss
  36. Gain
  37. Covalent
  38. Three
  39. Oxidation

40. Electronegativity

41. Covalent

42. Two

43. Acceptor

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44. Positive

45. Hydroxyl

46. Salt, neutralization

47. (i) blue (ii) red (iii) hydrogen

(IV) Basic, alkaline (v) solid

48. Potash alum ( $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$ )

49. Concentrated sulphuric acid

50. i. common salt      ii. Sugar

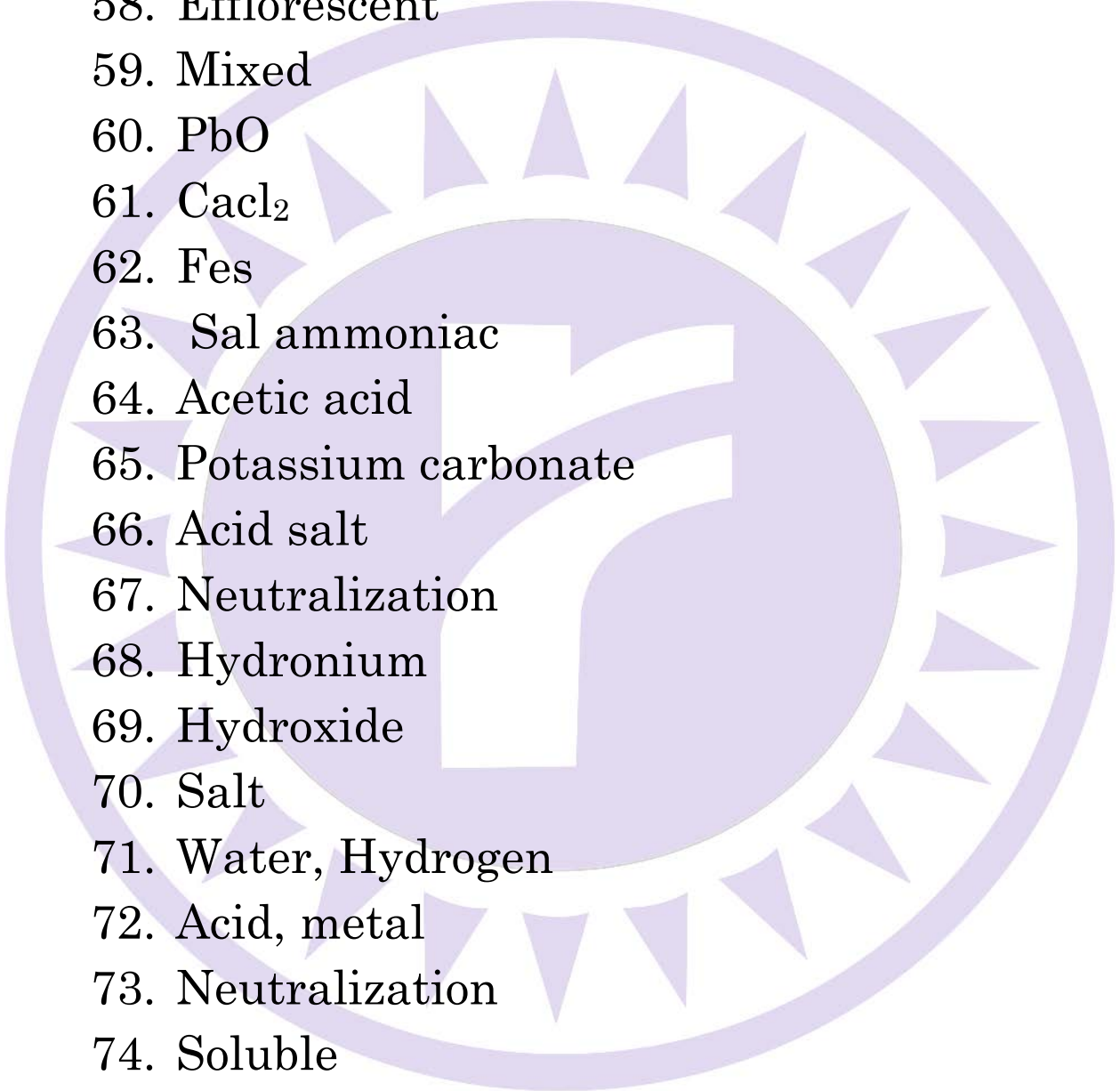
51. calcium sulphate

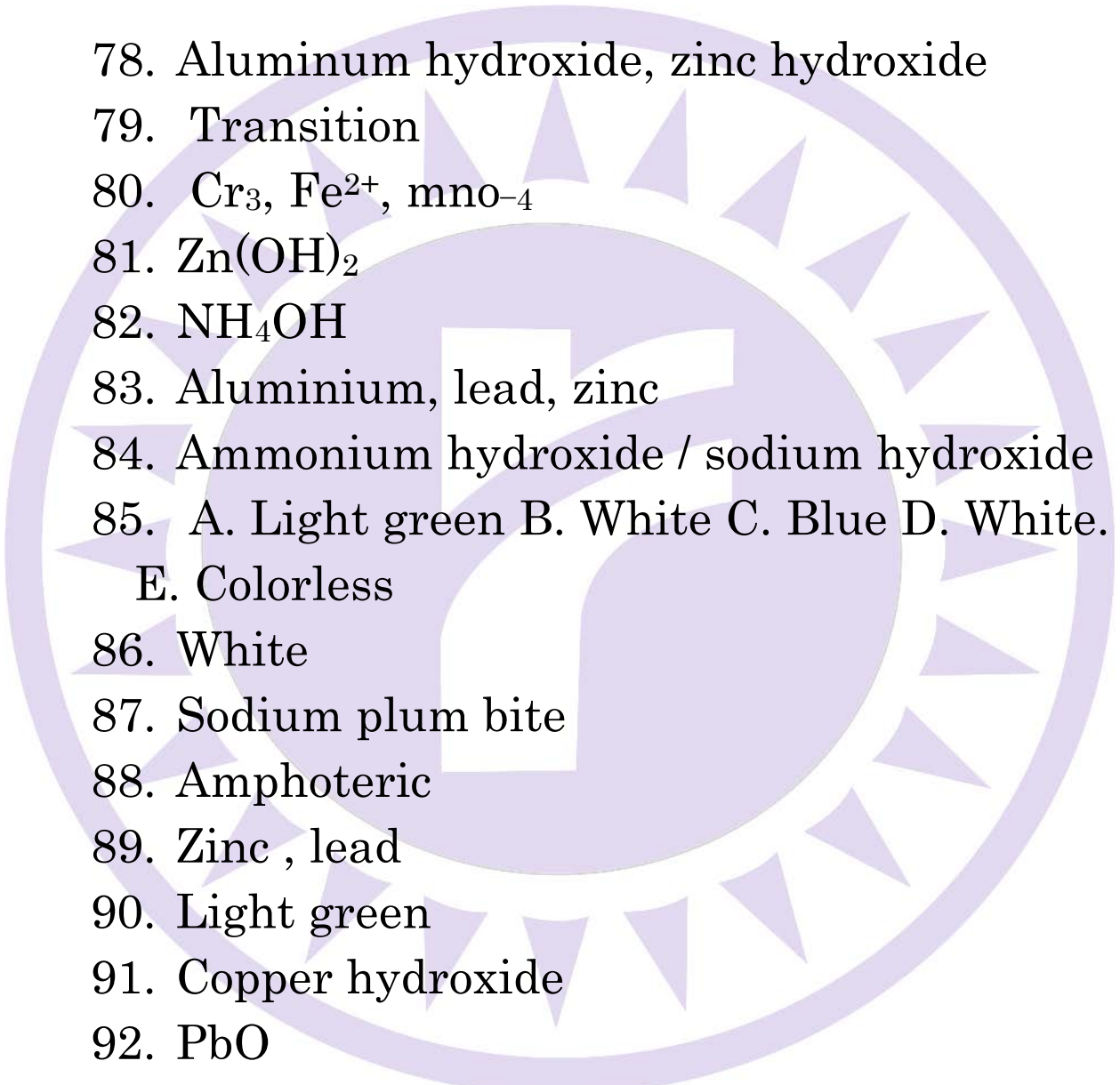
52. i. soluble salts      ii. Alkali

53. i.  $Mg(OH)_2$  [because all others are amphoteric hydroxides]

ii. Nitric acid [this is the only inorganic acid rest all are organic acids.]

54. Copper hydroxide

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55. Complex
  56. Phosphoric acid
  57. Deliquescent
  58. Efflorescent
  59. Mixed
  60. PbO
  61.  $\text{CaCl}_2$
  62. FeS
  63. Sal ammoniac
  64. Acetic acid
  65. Potassium carbonate
  66. Acid salt
  67. Neutralization
  68. Hydronium
  69. Hydroxide
  70. Salt
  71. Water, Hydrogen
  72. Acid, metal
  73. Neutralization
  74. Soluble
  75. An anhydrous
  76. Mixed salt
  77. Acidity

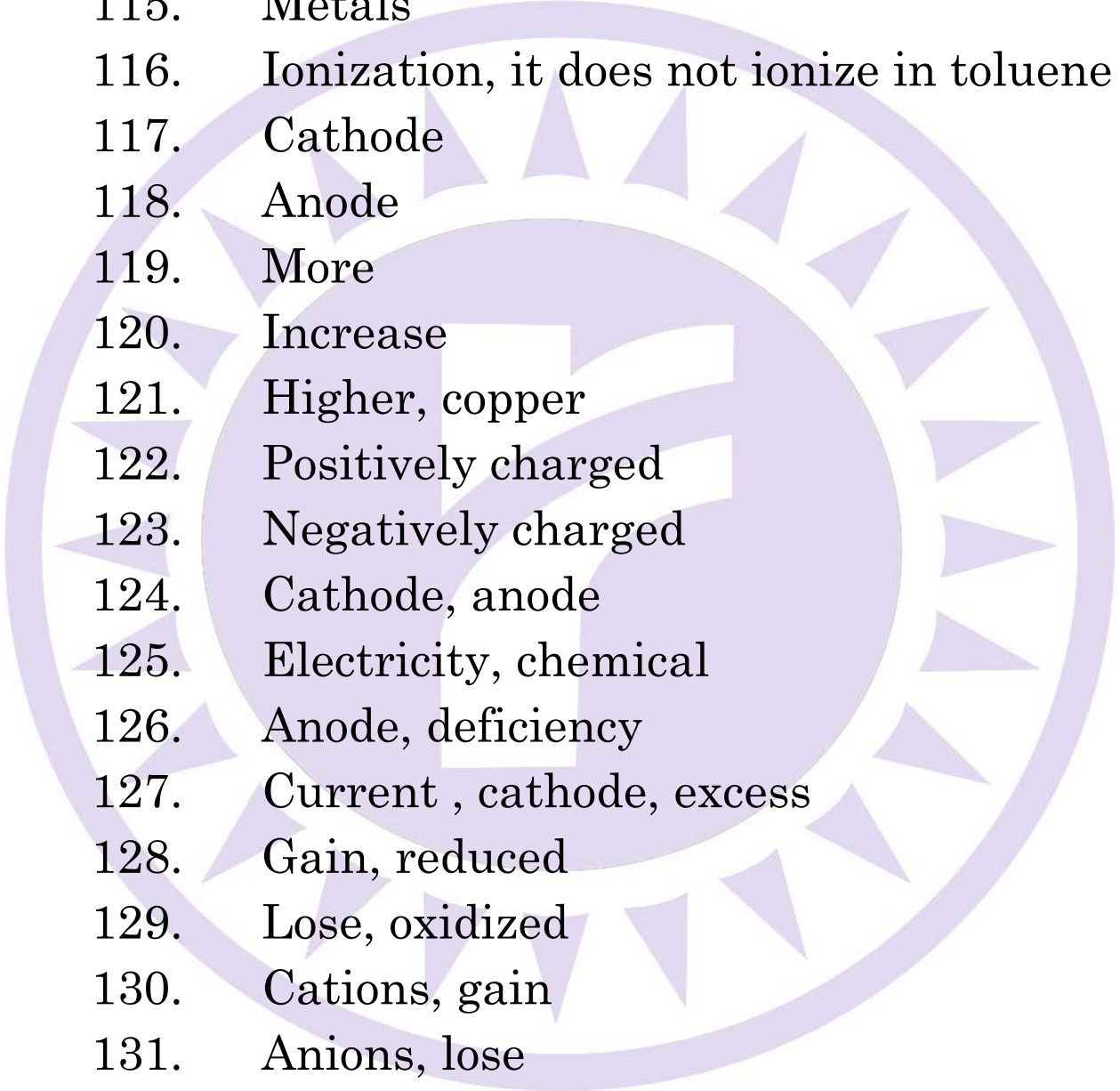
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78. Aluminum hydroxide, zinc hydroxide
  79. Transition
  80.  $\text{Cr}^{3+}$ ,  $\text{Fe}^{2+}$ ,  $\text{MnO}_4^-$
  81.  $\text{Zn}(\text{OH})_2$
  82.  $\text{NH}_4\text{OH}$
  83. Aluminium, lead, zinc
  84. Ammonium hydroxide / sodium hydroxide
  85. A. Light green B. White C. Blue D. White.  
E. Colorless
  86. White
  87. Sodium plum bite
  88. Amphoteric
  89. Zinc , lead
  90. Light green
  91. Copper hydroxide
  92.  $\text{PbO}$
  93. Ammonium hydroxide
  94.  $\text{PbSO}_4$
  95.  $(\text{NH}_4)_2\text{CO}_3$

96. Blue  
97.  $\text{Al(OH)}_3$   
98.  $\text{PbO}$   
99. Potassium carbonate  
100. Ammonia  
101. Lead chloride  
102. Two
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**5<sup>th</sup> chapter**

103. 36.072 %  
104. 62.9 %  
105. Molecule , atomocity  
106. Atmos , empirical formula
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107. Present in the molten state, present in the aqueous solution.  
108. Cathode  
109. Electrolyte, cations, anions,  
110. positively, cations  
111. negatively, anions.

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112. Ions
  113. Solid state
  114. Cannot
  115. Metals
  116. Ionization, it does not ionize in toluene
  117. Cathode
  118. Anode
  119. More
  120. Increase
  121. Higher, copper
  122. Positively charged
  123. Negatively charged
  124. Cathode, anode
  125. Electricity, chemical
  126. Anode, deficiency
  127. Current , cathode, excess
  128. Gain, reduced
  129. Lose, oxidized
  130. Cations, gain
  131. Anions, lose
  132. Reduced
  133. Higher
  134. Cathode, anode

135. i. Electrolyte ii. Nickel iii. Cathode iv.  
Anode v. cations
136. HCN and weak
137.  $\text{NiSO}_4$
138.  $\text{Na}[\text{Ag}(\text{CN})_2]$
139. HgO
140. Lead acetate
141. Chlorine
142. Sodium chloride
143. Carbon rods
144. Molecules , will not
145. Ions , lattice , electrodes
146. Melted
147. Cathode , anode
148. Electrolysis
149. Increases

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150. (a) E

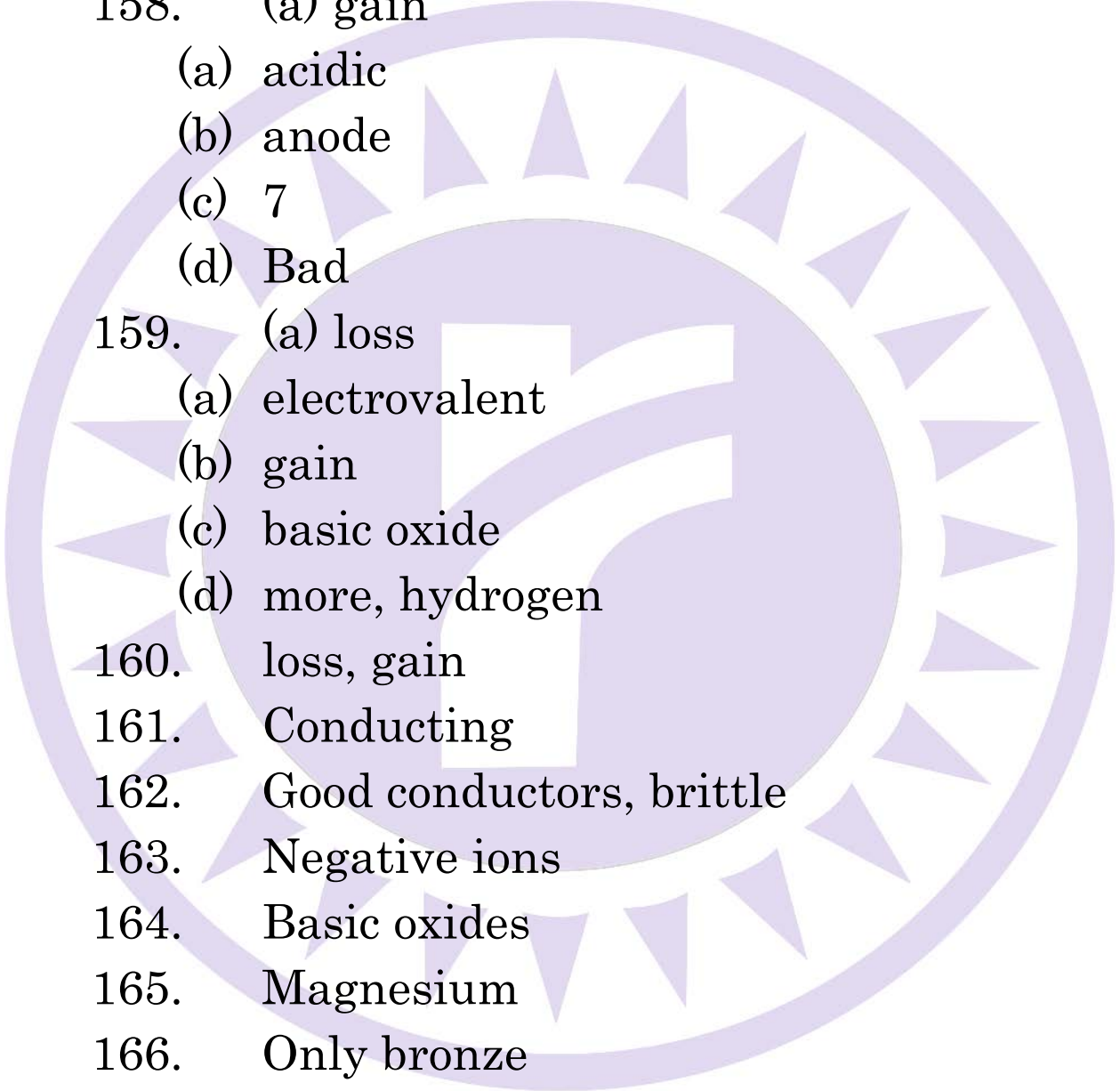
151. C

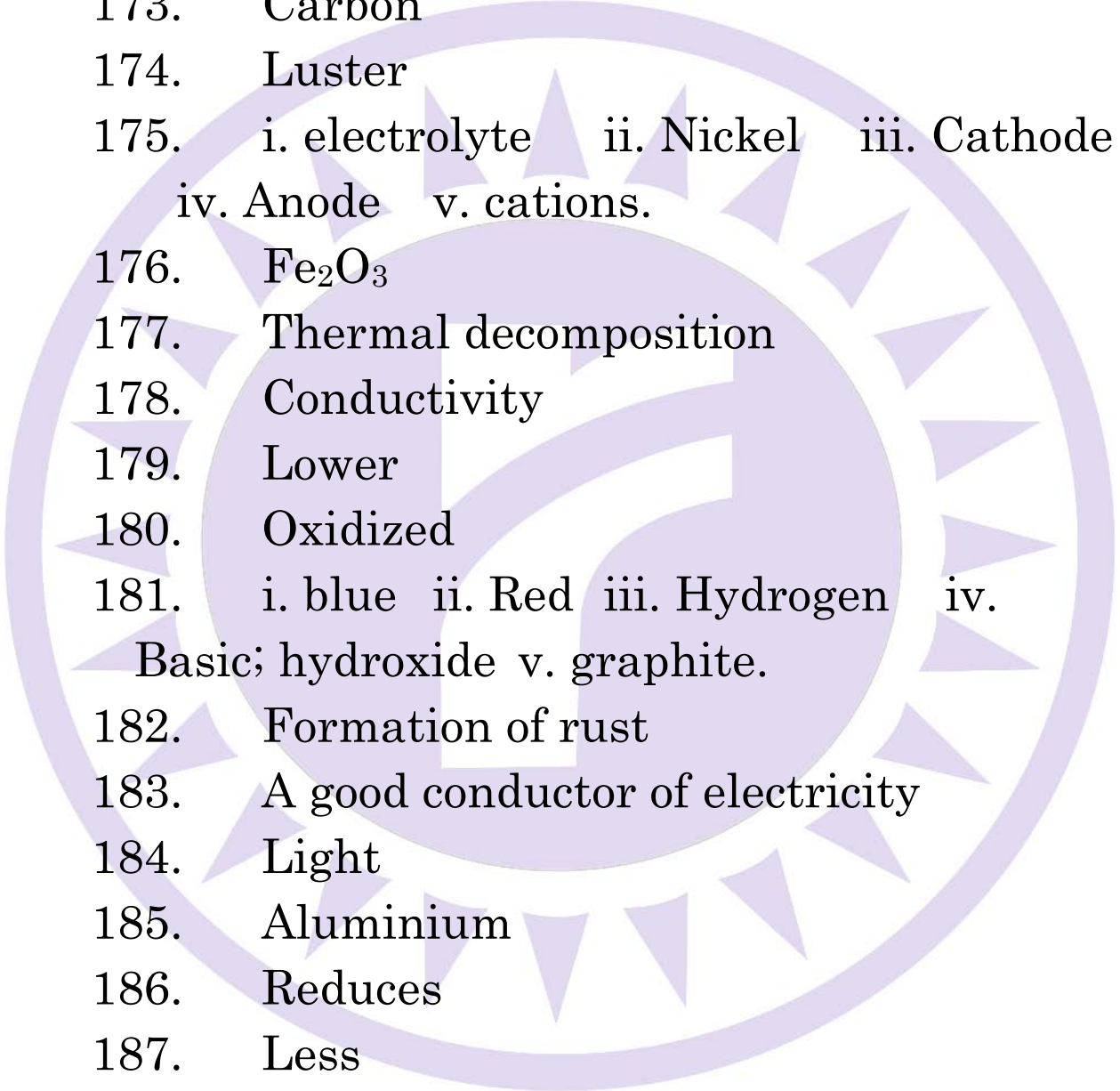
152.  $\text{Na}_3[\text{AlF}_6]$

153. Cu

154. Na

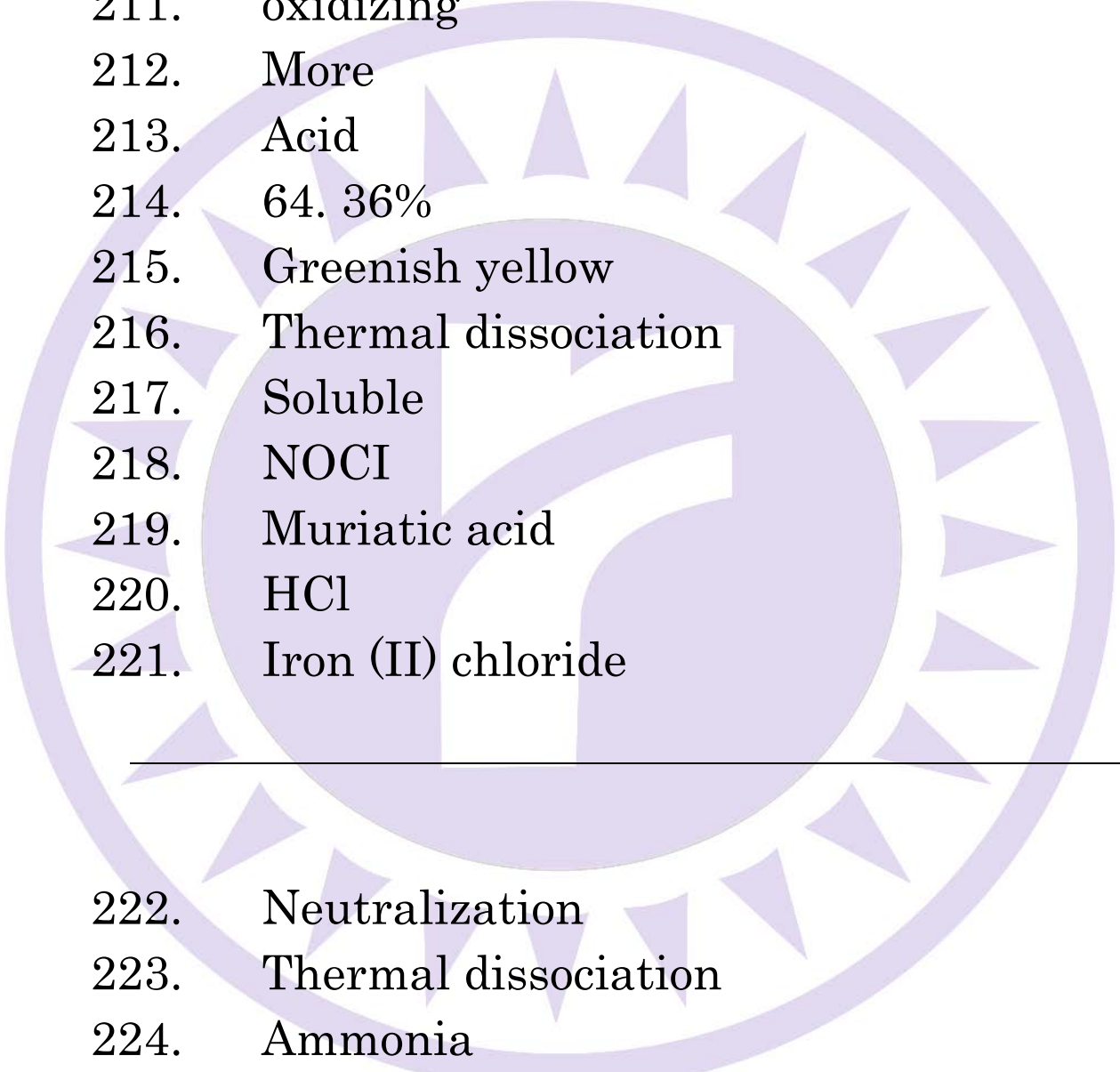


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155. Ag  
156. Mg  
157. Zn  
158. (a) gain  
    (a) acidic  
    (b) anode  
    (c) 7  
    (d) Bad  
159. (a) loss  
    (a) electrovalent  
    (b) gain  
    (c) basic oxide  
    (d) more, hydrogen  
160. loss, gain  
161. Conducting  
162. Good conductors, brittle  
163. Negative ions  
164. Basic oxides  
165. Magnesium  
166. Only bronze  
167. Cu-Zn-Ni  
168. Brass  
169. Bell metal

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170. Pb,Sn, and Sb  
171. Sn  
172. Copper alloy  
173. Carbon  
174. Luster  
175. i. electrolyte ii. Nickel iii. Cathode  
iv. Anode v. cations.  
176.  $\text{Fe}_2\text{O}_3$   
177. Thermal decomposition  
178. Conductivity  
179. Lower  
180. Oxidized  
181. i. blue ii. Red iii. Hydrogen iv.  
Basic; hydroxide v. graphite.  
182. Formation of rust  
183. A good conductor of electricity  
184. Light  
185. Aluminium  
186. Reduces  
187. Less  
188. Anode  
189. (a) steel (b) high  
190. Pig iron

191. i. amalgam ii. liquid iii. Semisolid  
iv. Dental amalgams v. reducing agent
192. i. Iodine [Because it is a sublimable substance] ii. Mercury [Because others are solids whereas mercury is liquid]
193. hematite is the chief ore of iron.
194. Limestone
195. coke
196. Iron III oxide
197. carbon monoxide
198. Constituents
199. Iron
200. Ferrous
201. Desired
202. Possessed

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203. CaO
204. Thermal dissociation
205. iron (II) chloride, simple displacement
206. Polar covalent compounds, contains
207. Silver nitrate, soluble, insoluble

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208. Iron (II) sulphide  
209. One, three, weight, oxidizes  
210. Lead(IV)oxide,  
211. oxidizing  
212. More  
213. Acid  
214. 64. 36%  
215. Greenish yellow  
216. Thermal dissociation  
217. Soluble  
218. NOCI  
219. Muriatic acid  
220. HCl  
221. Iron (II) chloride
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222. Neutralization  
223. Thermal dissociation  
224. Ammonia  
225. i. Ammonia ii. Alkaline iii.  
Ammonium iv. Hydroxyl v. dirty green  
226. alkalies

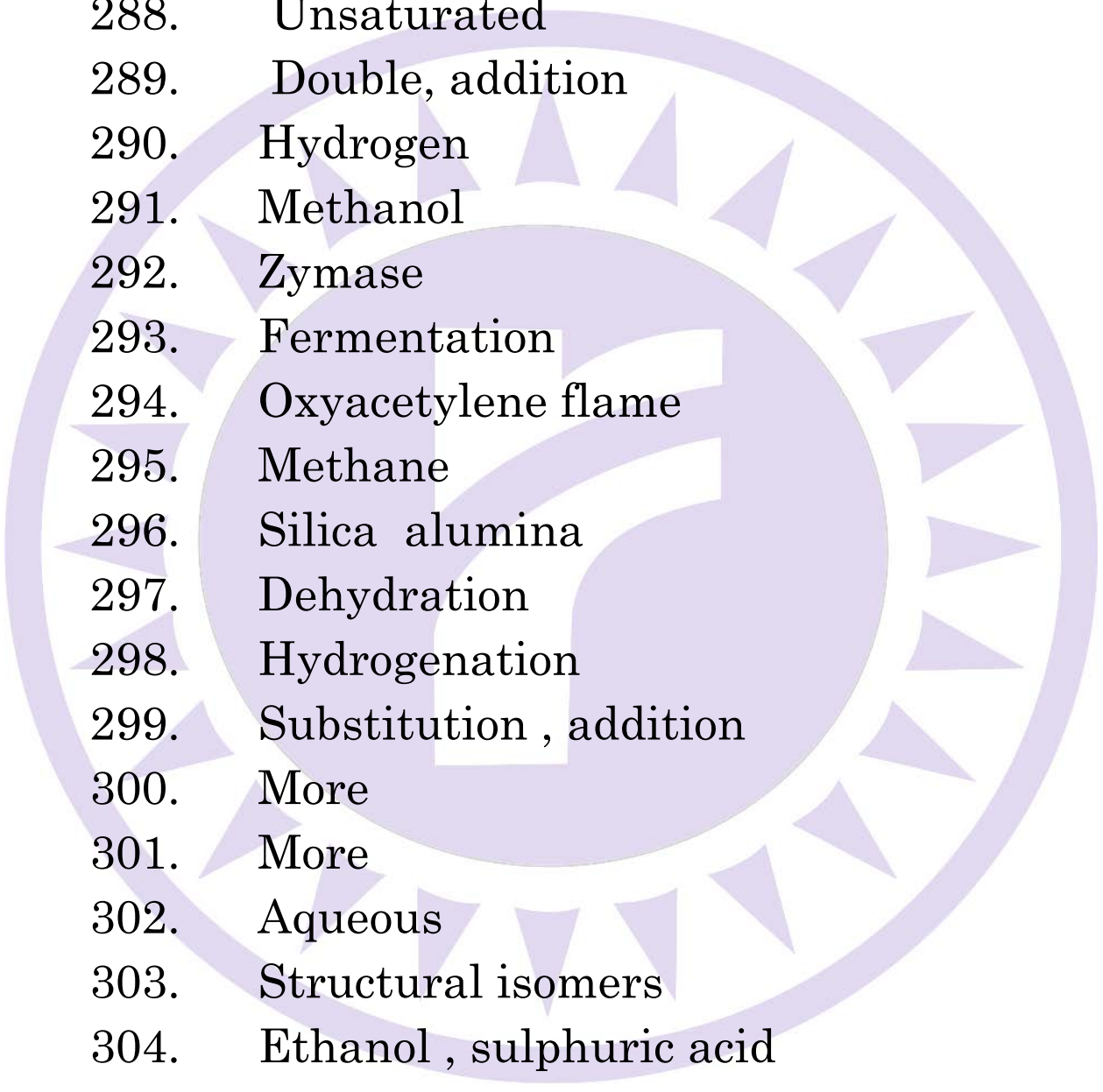
227.  $\text{Ca(OH)}_2$   
228. Air  
229. Promoter  
230. Reduces  
231. Yellow  
232. Brown  
233.  $\text{CaCl}_2 \cdot 8\text{NH}_3$   
234.  $\text{NH}_3$   
235. Nitrides  
236. Reducing
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237. 1:10, platinum gauge,  $800^\circ\text{C}$ , nitric oxide, oxygen, nitrogen dioxide, water  
238. Aqua fortis  
239. All glass  
240. Quartz  
241. Brown  
242.  $\text{FeSO}_4 \cdot \text{NO}$   
243. Yellow  
244. Sulphur  
245. Mn

246.  $\text{Cu}(\text{NO}_3)_2$   
247.  $\text{CaO} \cdot \text{Ca}(\text{NO}_3)_2$   
248. Concentrated  
249. Concentrated , Dilute

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250. Non-volatile  
251.  $\text{SO}_2$   
252. Bad  
253. Forward  
254. Sulphuric acid  
255. Washing  
256. Sulphur  
257. Trinitrotoluene  
258. Acidic  
259. Non-volatile  
260. Oxidizing  
261. C  
262.  $\text{C}_2\text{H}_2\text{O}_4$   
263. Less volatile  
264. Strong  
265. Hydronium

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266. i. homologous ii.  $C_nH_{2n+2}$
267. Unsaturated, Hydrocarbons,  
additions
268. Dehydration
269.  $H_2SO_4$
270. Hydrogenation
271. nickel
272. 5,15
273. Silica, alumina
274. Alkenes
275. Alkanes
276. Increase
277. Ethylene
278. propylene
279. Acetylene
280. Methane
281. Cycloalkene benzene
282. Nickel
283. Platinum , Palladium
284. Alkenes

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285. Alkynes
  286. Alkanes
  287. Homologous
  288. Unsaturated
  289. Double, addition
  290. Hydrogen
  291. Methanol
  292. Zymase
  293. Fermentation
  294. Oxyacetylene flame
  295. Methane
  296. Silica alumina
  297. Dehydration
  298. Hydrogenation
  299. Substitution , addition
  300. More
  301. More
  302. Aqueous
  303. Structural isomers
  304. Ethanol , sulphuric acid
  305. Hydrogen
  306. Water
  307. Hydrogen



- 308. Shorter
- 309. Ethanol
- 310. Methylated

\_\_\_\_\_ X                      X                      X                      X \_\_\_\_\_

