



ray (ICSE) classes

Objective Paper (Q)

Q1. Name the following:

1. Tendency of an element to form chains of identical atoms
2. Organic compounds having same molecular formula but different structural formula.
3. An example of saturated hydrocarbons
4. The type of reactions alkanes undergo.
5. The type of reactions alkenes undergo.
6. The gas commonly called as marsh gas.
7. General formula of alkanes
8. The solution which detects the presence of unsaturation in the given hydrocarbon.
9. The process by which ethane is obtained from ethanol.
10. The catalyst used during hydrogenation of alkene.
11. First organic compound synthesized.
12. Mixture of sodium hydroxide & Calcium oxide
13. The series in which members are arranged in the increasing order of their molecular masses.
14. Hydrocarbon used for artificial ripening of fruits
15. A reagent which is used to distinguish b/w ethene & ethyne.
16. General formula of alkenes
17. General formula of alkyne
18. Decomposition of higher molecular with hydrocarbon into a lower molecular with hydrocarbon.
19. Group derived from alkanes by loss of a hydrocarbon atom
20. System of assignment of names to organic compounds
21. General formula of aldehydes
22. An organic acid which is a chief constituents of vinegar.
23. Type of compound formed by the reaction between acetic acid and alcohol.
24. A compound which will give acetylene gas when treated with water.
25. Solid which can be used instead of concentrated sulphuric acid to prepare ethylene by the dehydration of ethanol.
26. Name a reagent which can be used to distinguish between ethane and ethene.
27. Name the addition product formed between ethene and water.

28. A gas acidic in nature which cannot be dried by passing through conc. H_2SO_4 .
29. Colour change of acidified $K_2Cr_2O_2$ soln. when SO_2 is passed through it.

30. The process by which SO_2 shows bleaching action.
31. Catalyst used in contact process.
32. Basicity of H_2SO_4 .
33. Solution that turns black when in contact with H_2S .
34. A neutral oxide of carbon when $\text{H}_2\text{SO}_4(\text{conc})$. Reacts with formic acid.
35. Atomicity of sulphur.
36. An organic acid reacts with H_2SO_4 . forms 2 oxides of C i.e Co and CO_2
37. Black substance formed when sugar is burnt.
38. Gas obtained on roasting Galena.
39. Number of acid salts H_2SO_4 has.
40. Gas obtained by dehydration of ethanol.
41. Gas with rotten egg smell.
42. Drying agent for SO_2 gas.
43. Anhydride of sulphuric acid.
44. A permanent bleaching agent.

45. Another name for nitric acid.
46. The undesirable compound formed when conc. H_2SO_4 is heated with NaNO_3
47. The gas giving nitric acid a brown colour.
48. The compound formed which causes yellowing of skin when HNO_3 fall on it.
49. A colourless neutral oxide if nitrogen.
50. The temp. at which conc. HNO_3 forms a constant boiling mixture.
51. The gas when hot and conc. HNO_3 falls on Cu.
52. The gas when cold and dilute HNO_3 falls on Cu.
53. The compound formed on which forms the brown ring.
54. A metal which turns passive with nitric acid.
55. Name the metals that react with nitric acid to liberate hydrogen.
56. Name the industrial method of preparation of nitric acid.
57. Nitric acid is one of the components of
58. Nitric acid consists of

59. The process by which ammonium is manufactured.
60. Drying agent for ammonia gas.
61. Metal nitride which on hydrolysis gives ammonia gas.
62. Catalyst used in Haber's process.
63. Two metallic oxides which are reduced by ammonia
64. Two solutions which react together to produce nitrogen.
65. Two gases which gives dense white fumes with ammonia.

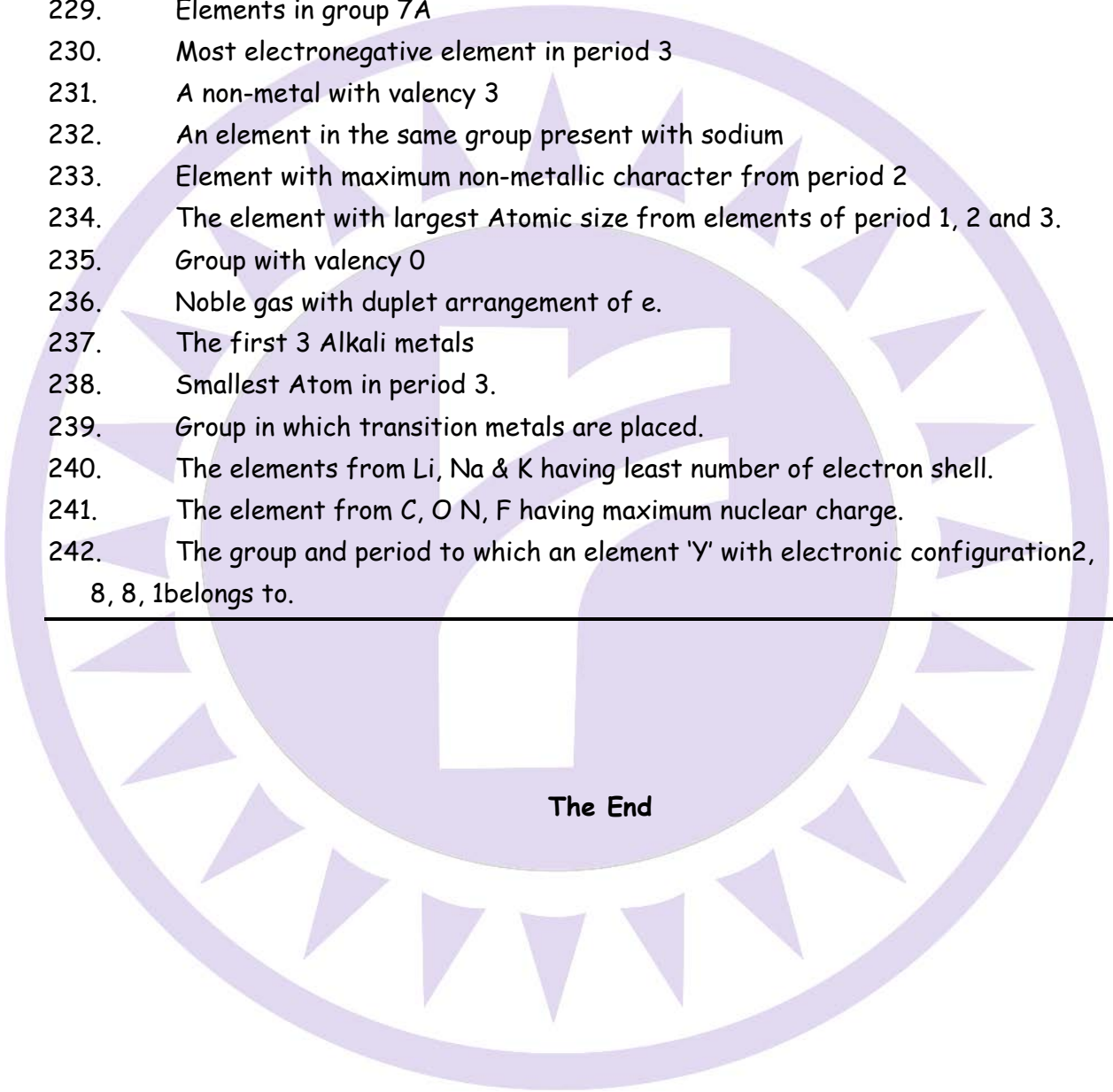
66. The solution which turns brown when it comes in contact with ammonia
 67. Experiment which demonstrate the extreme solubility of ammonia.
 68. A colourless gas which becomes reddish brown when it comes in contact with atmosphere(oxygen)
 69. Metal which directly combines with nitrogen on heating
 70. Aqueous solution of ammonia gas in water.
 71. Ammonia gas liquefied at low temperature
 72. An amphoteric oxide reduced by $\text{NH}_3(\text{g})$
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73. Name the two gases that combine to give a solid.
 74. Name the metals that do not react with dil.HCL acid to liberate H_2 .
 75. Name the two metals that react vigorously and explosively with dil.HCL acid.
 76. Common name for Hydrochloric acid.
 77. Common name for hydrogen chloride
 78. Acidic constituents of gastric juices.
 79. Two volatile acids
 80. A non-volatile acid
 81. An undesirable compound formed into the preparation of HCl above 200°C
 82. An insoluble chloride
 83. Two colourless gases which combine to form a solid.
 84. A solution used to test HCl acid
 85. A chloride soluble in hot water & insoluble in cold water
 86. The solvent for silver chloride
 87. The solvent for gold & platinum
 88. A gas which gives dense white fumes with hydrochloric acid
 89. A metal insoluble in mineral acid but soluble in aqua regia
 90. A gas which is neutral to dry litmus but turn moist blue litmus red
 91. Salt of a weak acid
 92. A weak acid
 93. A trivalent metal chloride
 94. A metal which liberates H_2 explosively
 95. A basic gas, NH_3 , reacts with an acidic gas, HCl, to form a white solid
 96. A gaseous compound-valent compound which dissolves in water & conducts electricity
 97. Two metals which do not displace H_2 from acid.
 98. Two metals which displace H_2 from acid
 99. A gas which is non-combustible & nor a supporter of combustion

100. A blue green chloride.
101. Name one lead compound that can be used to oxidize hydrogen chloride to chlorine.
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102. Allotropic modification of carbon and good conductor of electricity
103. Sulphide ore of mercury
104. Bonding present in metallic chloride
105. The process of removal of gangue from ore
106. A naturally occurring compound of metal from which metal is extracted cheaply, profitably, and conveniently
107. The process by which sulphide ores are concentrated.
108. The process of heating of ore in the presence of air
109. The process by which impurities are removed from iron obtained from blast furnace
110. The most common ore of aluminum
111. Metallic oxides which are reduced by aluminum
112. The formula of slag
113. The chemical name of slag
114. The metals added to steel to make it stainless steel
115. Metal which is rendered passive on reaction with conc. Nitric acid.
116. Gas obtained when zinc blende is roasted
117. The process of coating thin layers of Zn over the surface of iron.
118. Highly electropositive metals.
119. Least electropositive metals
120. A metal which can be reduced easily by thermal decomposition of its oxide.
121. A metallic oxide most difficult to reduce
122. A metal which can form +2 and +3 ions
123. A metal which can be reduced thermally from its oxide
124. Two impurities in pig iron other than carbon, silicon
125. The form of iron which has 1 -5% carbon impurity and is used in making nuts and bolts.
126. Two metals rendered passive or unreactive by conc. Or fuming nitric acid.
127. A metal other than mercury present in liquid amalgam
128. Name the ore of zinc containing its sulphide.
129. Name the substance used to reduce the roasted ore.
130. Name the oxide which acts as flux in the blast furnace
131. Name an allotrope of a non-metal that allows electricity to pass throughout.
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132. Name a non-metal that has a metallic luster and sublimates on heating
133. Name the important commercial forms of iron.
134. Name the three methods of obtaining a metal from its ore.
135. Metal which is liquid at room temperature
136. Metal which is neither malleable nor ductile
137. Metal which is soft and heavy
138. Metal which can be cut with a knife
139. Hard but lacks mechanical strength.
140. Is brittle at ordinary temperature.
141. Forms a liquid alloy at ordinary temperature.
142. Is soft and floats on water
143. Is divalent and forms an unstable hydride
144. Is trivalent and forms an acidic oxide
145. Is trivalent and forms an amphoteric oxide
146. Metal with low boiling point
147. Forms an amalgam
148. Is difficult to extract from its ore
149. Has a low melting point.
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150. Compounds which conduct electricity when dissolved in water in the molten state.
151. The decompositions of a chemical compound in the aqueous or fused state by the passage of direct electric current.
152. Electrode connected to positive terminal of the battery
153. Electrode connected to negative terminal of the battery
154. Compounds which do not conduct electricity when dissolved in aqueous or molten state.
155. Electrolytes which almost completely dissociate in fused or aqueous state.
156. Electrolytes which are partially dissociated in fused or aqueous state.
157. It is the vessel in which electrolyte is carried out.
158. Ions which migrate to anode
159. Ions which migrate to cathode
160. The number of positive charges equals the number of negative charges in the electrolytic solution.
161. The process due to which an ionic compound in the fused or in aqueous state dissociates into ions by passage of electric current.

162. An ionic compound, when added to water, it dissociates.
163. A polar solvent
164. A metal which ionize most readily.
165. A metal which ionize least readily.
166. It is the preferential discharge of ions present in an electrolyte at respective.
167. An inert electrode
168. An active electrode.
169. Product formed at anode during electrolysis of molten lead bromide.
170. Acid used for acidification of pure water during electrolysis.
171. Product formed at cathode during the electrolysis of acidified water.
172. Product formed at anode during the electrolysis of acidified water.
173. Product formed at anode during the electrolysis of aqueous copper sulphate inert electrodes.
174. Product formed at anode during the electrolysis of aqueous copper sulphate using active copper electrodes.
175. The electrolytic process of decomposition of superior metal on the surface of baser metal.
176. The article to be plated is placed at the electrode.
177. The metal to be plated on the article is always placed at the electrode.
178. Anode used during electroplating of an article with nickel.
179. The electrode which diminishes in mass during electroplating.
180. The process by which metals containing impurities are purified electrolytically to give a pure metal.
181. Metals generally refined by electrolysis.
182. The process of extraction of the metals by electrolysis.
183. Metals which are generally extracted by electrolysis.
184. Metals that can be extracted from their oxides by the thermal decomposition.
185. Substances which do not allow an electric current to flow through them.
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186. A basic solution which does not contain a metallic element.
187. An alkali which on dissociation produces a high concentration of hydroxyl ions.
188. A complex salt solution used for testing a basic gas lighter than air.
189. A base which reacts with hydrochloric acid to give a salt which on hydrolysis gives a slight acidic solution.

190. An ion which combines with a polar covalent molecule to form ammonium ion.
191. The ion other than ammonium ion formed when ammonia dissolves in water.
192. Two bases which are not alkalies.
193. A normal salt and an acid salt of the same acid.
194. A salt insoluble in cold water but soluble in hot water.
195. A substance which changes the blue colour of copper sulphate crystals to white.
196. Two crystalline substances which don't contain water crystallisation.
197. The name and formula of the acid salt which gives sodium ions and sulphate ions in solution.
198. Strong acid containing chlorine.
199. Two dibasic acids containing sulphur.
200. Acid anhydride of sulphuric acid.
201. Two monobasic acids containing nitrogen.
202. The name and formula of acid salt which gives sodium ions and sulphate ions.
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203. Force which holds 2 or more atoms together in a stable molecule.
204. An atom that has a charge.
205. Type of reaction which results in loss of electrons.
206. Conversion of Fe^{3+} to Fe^{2+} .
207. Bonding between 2 atoms due to transfer of electrons
208. Pair of electrons not shared with any other atom.
209. Electrons that form co-ordinate bond.
210. Examples of polar covalent compound.
211. Examples of a non-polar covalent compound.
212. Type of bonding between a metal and a non-metal.
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213. The valency of elements of group 2.
214. The alkali metal in period
215. The liquid Halogen
216. The metalloid in period
217. The inert gas having configuration 2,8
218. The element with largest atomic size in period 2
219. The element with highest ionisation potential from the elements of period 3.
220. The strongest oxidizing agent in period 3.
221. The formula of the oxide of element in period 3 group IA Sodium oxide.
222. The element with highest electronegativity from the elements Cl, S,P,Si.
223. 2 elements which are in the same group as carbon

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224. Alkali metal in period 3
225. Alkaline earth metal in period 3.
226. Two metals in period 2
227. Two metals in period 3
228. Number of shells in 3rd period
229. Elements in group 7A
230. Most electronegative element in period 3
231. A non-metal with valency 3
232. An element in the same group present with sodium
233. Element with maximum non-metallic character from period 2
234. The element with largest Atomic size from elements of period 1, 2 and 3.
235. Group with valency 0
236. Noble gas with duplet arrangement of e.
237. The first 3 Alkali metals
238. Smallest Atom in period 3.
239. Group in which transition metals are placed.
240. The elements from Li, Na & K having least number of electron shell.
241. The element from C, O N, F having maximum nuclear charge.
242. The group and period to which an element 'Y' with electronic configuration 2, 8, 8, 1 belongs to.
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The End