



CHAPTER-1

HYDROGEN CHLORIDE & HYDROCHLORIC ACID

- (a) A white ppt. is formed which is soluble in excess of ammonium hydroxide solution.
 - (b) Dense white fumes are seen.
 - (c) A white ppt. is formed which gets dissolved on heating.
 - (d) It dissolves.
 - (e) Black colour CuO dissolves in it to give a blue colour solution.
 - (f) Gives a greenish yellow gas.
2. Dense white fumes of ammonium chloride are seen.
3. Lead (IV) oxide oxidizes hydrochloric acid and chlorine gas is liberated.
4. The blue coloured litmus solution in water when enters the round bottomed flask turns red.
5. Moist blue litmus paper initially turns red and then gets bleached.
6. Greenish yellow colour of chlorine water fades away.
7. Dense white fumes of ammonia are seen.

CHAPTER-2 AMMONIA

1. (a) Red litmus soln turns blue.

(b) it burns with greenish yellow flame forming nitrogen and water.

(c) The colourless nitric oxide undergoes further oxidation, to give redish brown fumes of nitrogen dioxide.

The platinum continues to glow even after the heating is discontinued.

(d) Black coloured copper oxide changes to pinkish red metal, a colourless liquid is obtained which turns anhydrous copper sulphate blue

(e) Buff yellow lead (II) oxide is reduced to greyish metallic lead, a colourless liquid is obtained which turns anhydrous copper sulphate blue.

(f) Dense white fumes of ammonium chloride are seen.

(g) Neutral litmus (purple) turns blue in alkaline solution.

2. A white precipitate is observed which remains insoluble in excess of ammonia.

3. Ammonia

CHAPTER-3 NITRIC ACID

1. Brown vapours are seen in the bottle and the nitric acid turn yellowish in colour.
2. A redish brown gas NO_2 is liberated.
3. The yellow colour in the concentrated nitric acid bottle appears after sometime due to the thermal decomposition of concentrated nitric acid in to redish brown nitrogen dioxide which gets dissolved in nitric acid and imparts it yellow colour.

CHAPTER-4 SULPHURIC ACID

1. (a)

- Black compound dissolves
- Light green solution formed
- Gas with rotten egg smell is given out

(b)

- Metal dissolves with dil. H_2SO_4
- Gas liberated which burns with pop siund.

(c) Yellow particles of S are deposited in the tube.

(d) Pink colour solution burns colourless.

(e)

- Pink colour solution is decolourised.
- Yellow S particles are seen deposited in the test tube.
- Solution is turbid (not turns parent)

(f)

- A white ppt is observed.
- Ppt insoluble in all mineral acids.

(g) Volume of the acid is increased after sometime.

(h) The water spurts out.

(i)

- Green colour disappears.
- Crystals turns white amorphous.

(j)

- Ribbon continues to burn with a dazzling white light.
- Yellow particles of S are deposited on the jar.

2. Copper (II) oxide



CHAPTER-5 ORGANIC CHEMISTRY

- (a) Brown colour of bromine gets discharged.
 - (b) A red precipitate of copper acetylide is formed.
 - (c) A white precipitate of silver acetylide is formed.
 - (d) Brisk effervescence of colourless, odourless gas are seen which burns with popping sound.
 - (e) The purple colour of potassium permanganate decolourises.
2. akjzdfbmnergrknv page no. 132